

Binary Ionic Compounds

- Consist of only _____ elements.
- Name the _____ ion (the _____).
- Name the _____ ion (the _____), changing the ending to _____.
- When metals that can form _____ than one type of _____ are in a compound, use a _____ in parentheses after the name of the _____ to show the _____.

Examples: NaCl _____

MgO _____

Cu₂S _____

SnCl₄ _____

Ternary Ionic Compounds

- Made up of _____ elements.
- Name the _____ then name the _____ without changing the ending to "ide".

Examples: Na₂SO₄ _____

FeCrO₄ _____

Naming Molecular Compounds

- The elements are named in the _____ they appear in the _____.
- _____ are used to denote the _____ of atoms of each _____ in the molecule. An exception is that the _____ element named is given a _____ only if there is more than _____ atom of that element in the _____.
- The "o" or "a" at the _____ of a _____ is _____ when the word following the _____ begins with a _____.
- The _____ element's ending is changed to _____.

Examples: ICl_3 _____
 As_2O_5 _____

Naming Hydrocarbons:

Alkane: $\text{C}_n\text{H}_{2n+2}$

Alkene: C_nH_{2n}

Examples: C_4H_{10} _____
 C_2H_4 _____

The Chemistry Quiz

CR1. _____

CR2. _____

1. _____

2. _____

3. _____

4. _____

5. _____